

Portal

GENERAL CHEMISTRY II / جميع الشعب

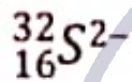
Question 1

Not yet answered

Marked out of 1.5

Flag question

What is the number of protons, electrons and neutrons in:



Select one:

- a. 16 protons, 18 electrons and 16 neutrons
- b. 14 protons, 16 electrons and 32 neutrons
- c. 16 protons, 14 electrons and 16 neutrons
- d. 16 protons, 18 electrons and 32 neutrons
- e. 14 protons, 16 electrons and 16 neutrons

Question 2

Not yet answered

Marked out of 5

Flag question

Evaluate the following expression to the correct number of significant figures.

$$4,2680 \times 0.82057 \times 373.150$$

$$(744.00/76.000) \times 2.6880$$



Select one:

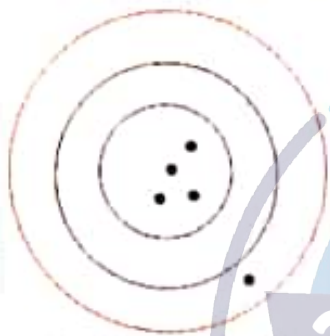
- a. 50.
- b. 49.6632
- c. 49.663
- d. 50.000
- e. 49.7

Question 3

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out of

Question

Describe the following set of experiments in terms of precision and accuracy



Select one:

- a. The experiment is both accurate and precise
- b. Non of the statements is correct
- c. The experiment is accurate but not precise
- d. The experiment is precise but not accurate
- e. The experiment is neither accurate nor precise

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Question 4

Not yet answered

Marked out of 1.5

Flag question

What is the name of the compound, P_4Te_7 ?

Te = Tellurium

Select one:

- a. phosphorus(IV) telluride
- b. tetraphosphorus heptatelluride
- c. phosphorus telluride
- d. phosphorus(VII) telluride
- e. tetraphosphorus telluride

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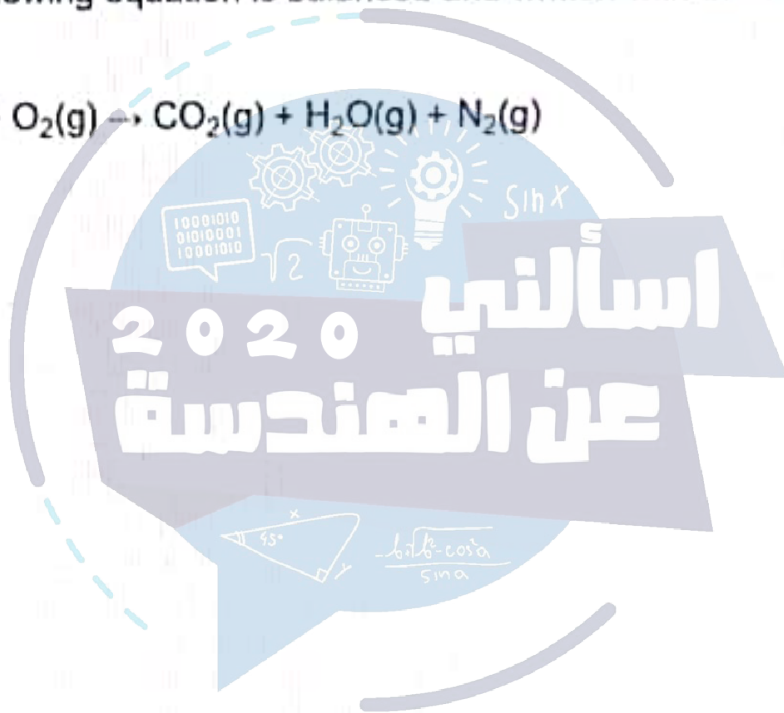
out of

When the following equation is balanced and written with the smallest whole number coefficients, what is the coefficient of H₂O?



Select one:

- a. 3
- b. 16
- c. 10
- d. 1
- e. 6



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GENERAL CHEMISTRY II / جميع الشعب

The SI base units of temperature and mass, respectively, are:

Select one:

- a. Kelvin and kilogram
- b. Celsius and milligram
- c. Kelvin and gram
- d. Fahrenheit and kilogram
- e. Fahrenheit and gram



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Question 7

Not yet answered
Marked out of 1.00

Flag question

An example of an element is

Select one:

- a. Table salt, NaCl
- b. Glucose, $C_6H_{12}O_6$
- c. Helium, He
- d. An oxide of iron, Fe_3O_4
- e. Limestone, $CaCO_3$

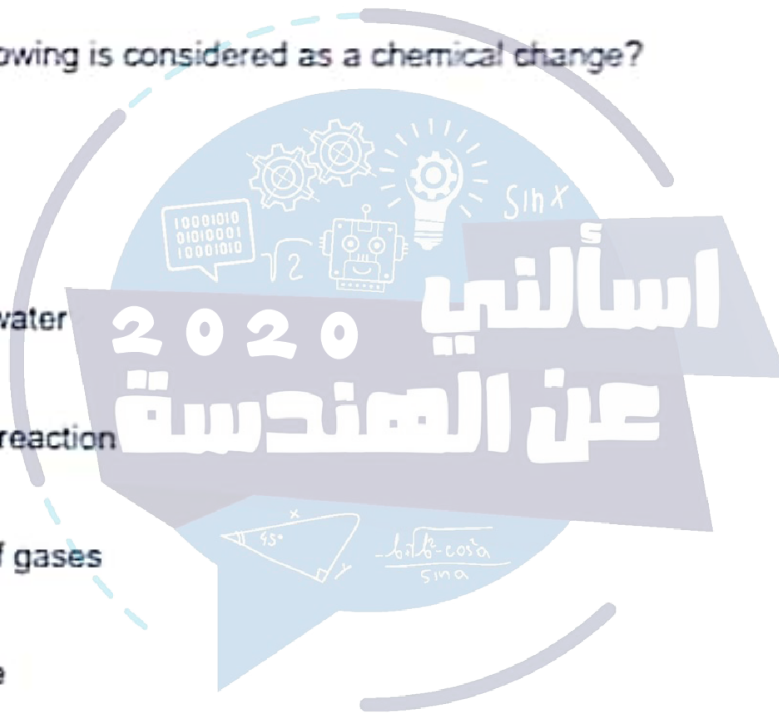
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GENERAL CHEMISTRY II / جميع الشعب

Which of the following is considered as a chemical change?

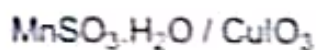
Select one:

- a. Freezing of water
- b. Combustion reaction
- c. Deposition of gases
- d. Melting of ice
- e. Sublimation of iodine



GENERAL CHEMISTRY II / جميع الشعب

Name the following compounds:



Select one:

- a. Manganese(II) Sulfate monohydrate / Copper(I) periodate
- b. Manganese(II) sulfide hydrate / Copper(I) iodate
- c. Manganese(II) thiosulfate monohydrate / Copper(I) iodide
- d. Manganese(II) sulfite monohydrate / Copper(I) iodate
- e. Magnesium Sulfate monohydrate / Copper(I) iodide

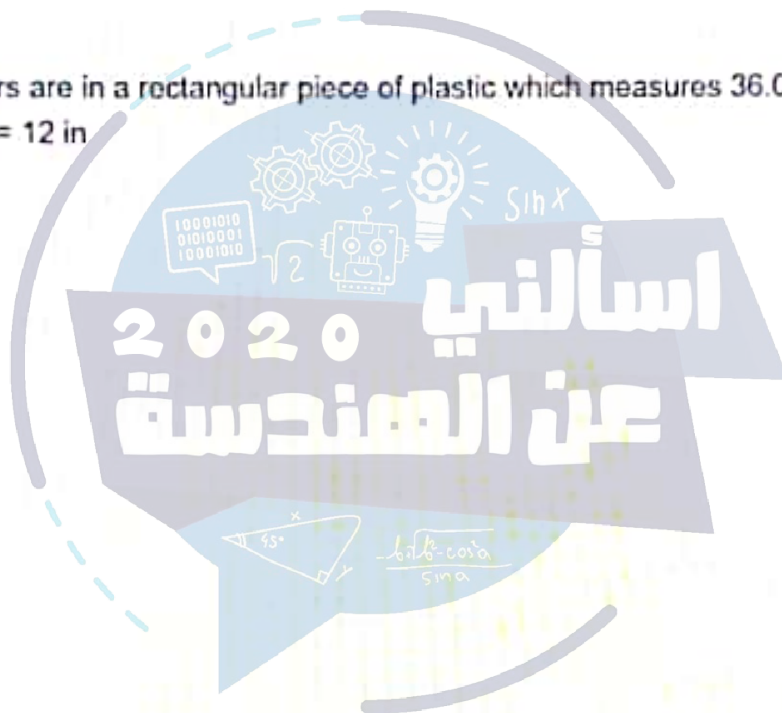
GENERAL CHEMISTRY II / جميع الشعب

How many square meters are in a rectangular piece of plastic which measures 36.0 feet by 20.0 feet?

1 m = 39.37 in., 1 ft. = 12 in

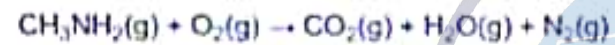
Select one:

- a. 66.9 m²
- b. 37.0 m²
- c. 60.7 m²
- d. 44.6 m²
- e. 20.8 m²



GENERAL CHEMISTRY II / جميع الشعب

When the following equation is balanced and written with the smallest whole number coefficients, what is the coefficient of H_2O ?



Select one:

- a. 3
- b. 16
- c. 10
- d. 1
- e. 6

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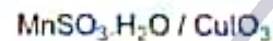
Question 9

Not yet answered

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Name the following compounds:



Select one:

- a. Manganese(II) Sulfate monohydrate / Copper(I) periodate
- b. Manganese(II) sulfide hydrate/ Copper(I) iodate
- c. Manganese(II) thiosulfate monohydrate / Copper(I) iodide
- d. Manganese(II) sulfite monohydrate / Copper(I) iodate
- e. Magnesium Sulfate monohydrate / Copper(I) iodide

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GENERAL CHEMISTRY II / جميع الشعب

Question 10

Not yet answered

Marked out of

Flag question

How many square meters are in a rectangular piece of plastic which measures 36.0 feet by 20.0 feet?

1 m = 39.37 in., 1 ft. = 12 in

Select one:

- a. 66.9 m²
- b. 37.0 m²
- c. 60.7 m²
- d. 44.6 m²
- e. 20.8 m²

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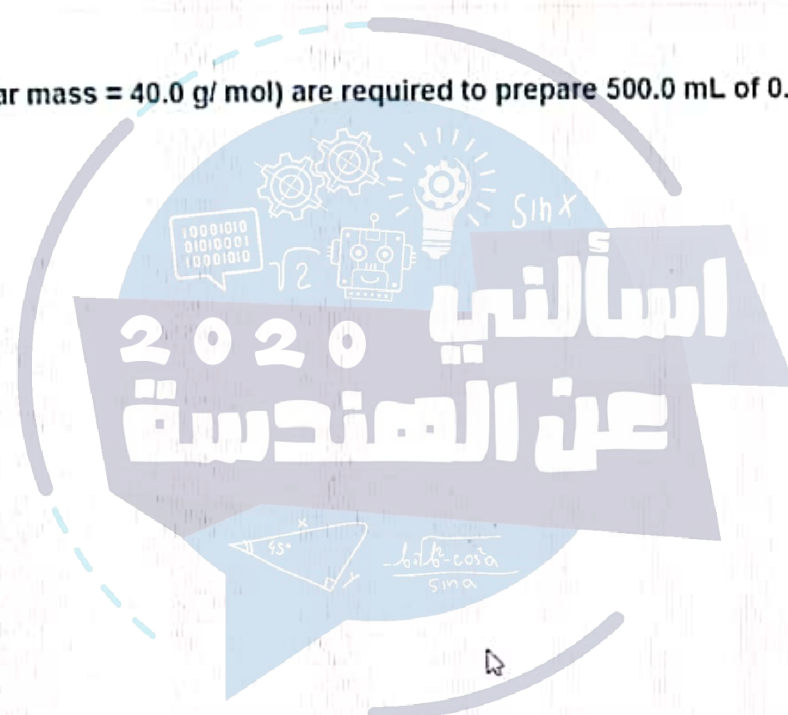
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AL CHEMISTRY II / جميع الشعب

How many grams of NaOH (molar mass = 40.0 g/ mol) are required to prepare 500.0 mL of 0.150 M solution?

Select one:

- a. 3.00
- b. 0.60
- c. 2.40
- d. 1.20
- e. 1.80



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GENERAL CHEMISTRY I / جميع الشعب

What is the oxidation number of Mn in MnO_4^- ?

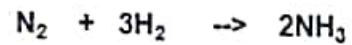
Select one:

- a. +4
- b. +3
- c. +6
- d. +5
- e. +7



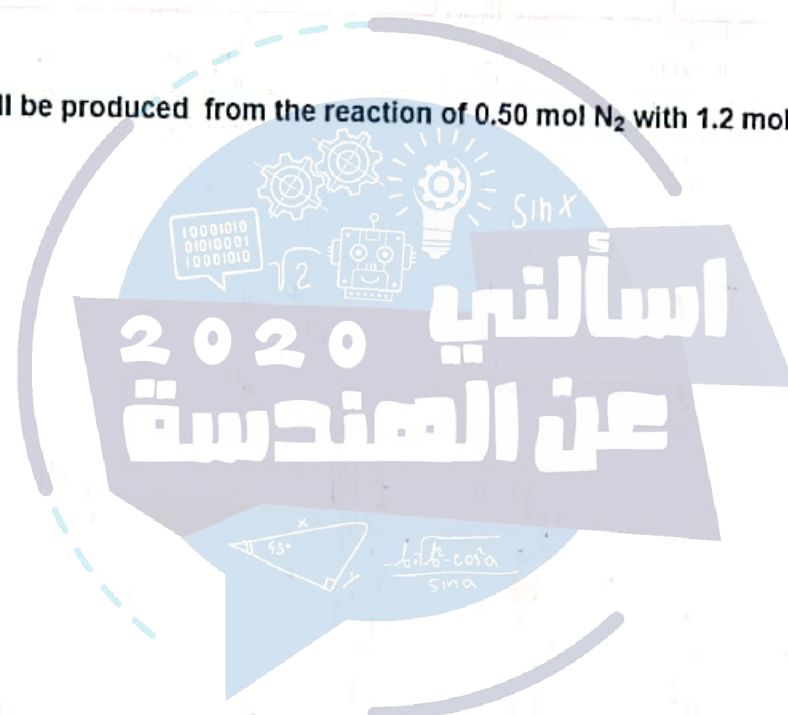
L CHEMISTRY II / جميع الشعب

How many moles of NH_3 will be produced from the reaction of 0.50 mol N_2 with 1.2 mol of H_2 according to the following equation?



Select one:

- a. 1.67
- b. 0.80
- c. 1.0
- d. 0.67
- e. 1.33



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AL CHEMISTRY II / جميع الشعب

A pure compound composed of C, H and O has a molar mass of 90.09 g/ mol. It gives the following elemental mass percentages:

C= 40.0% and H= 6.71%. What is the molecular formula of the compound?

Molar mass (g/mol) : C = 12.01; H = 1.008; O = 16.00

Select one:

- a. $C_2H_4O_2$
- b. $C_3H_6O_3$
- c. $C_4H_8O_4$
- d. $C_5H_{10}O_5$
- e. $C_6H_{12}O_6$

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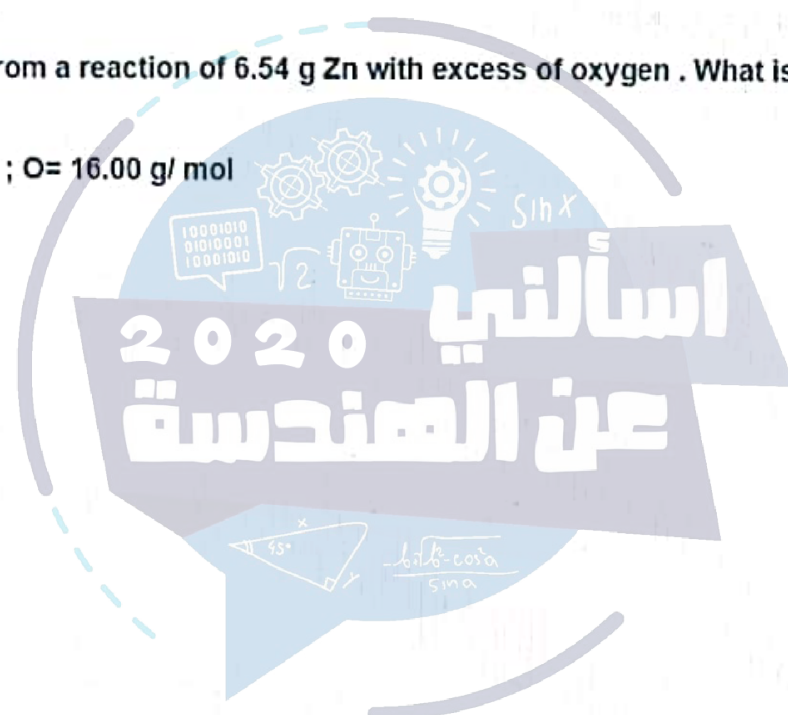
L CHEMISTRY II / جميع الشعب

If 7.32 g of ZnO are obtained from a reaction of 6.54 g Zn with excess of oxygen . What is the percent yield of the reaction?

Molar mass : Zn = 65.38 g/ mol ; O= 16.00 g/ mol

Select one:

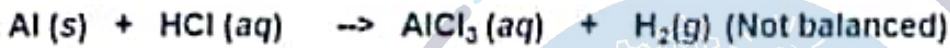
- a. 65%
- b. 75%
- c. 70%
- d. 90%
- e. 85%



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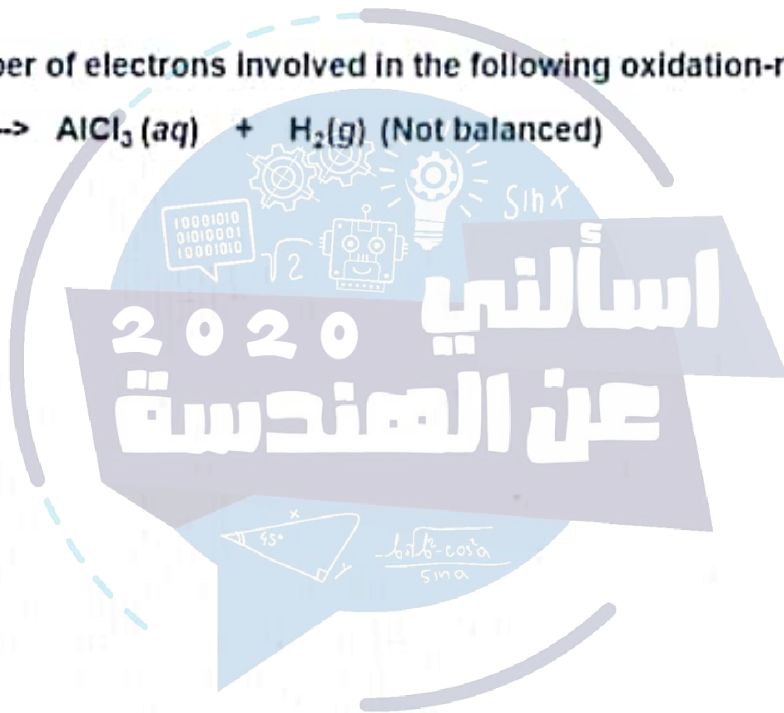
GENERAL CHEMISTRY II / جميع الشعب

What is the total number of electrons involved in the following oxidation-reduction reaction?



Select one:

- a. 6
- b. 2
- c. 8
- d. 10
- e. 4

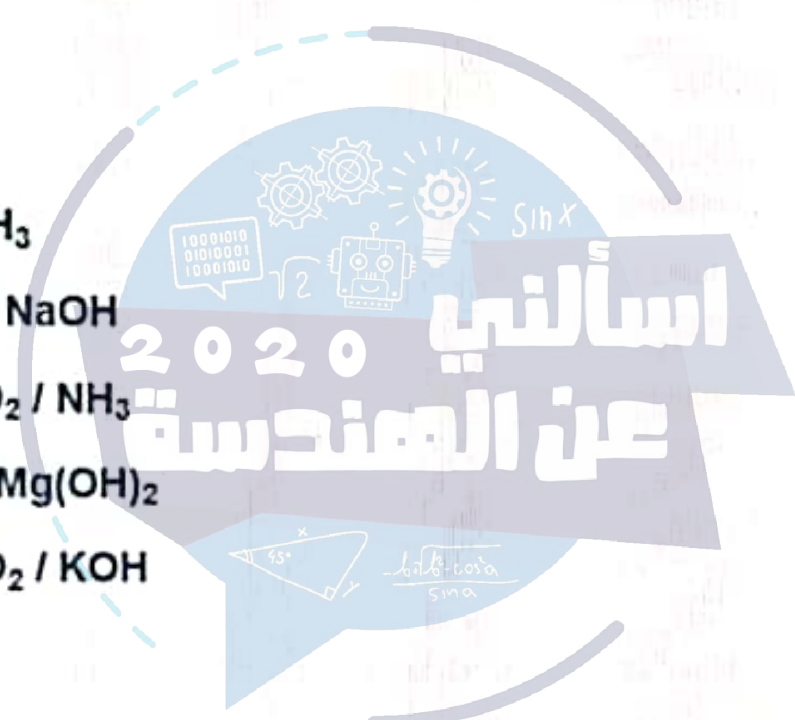


GENERAL CHEMISTRY I / جميع الشعب

Which of the following pairs is: weak acid / strong base ?

Select one:

- a. HCl / NH_3
- b. $\text{HClO}_4 / \text{NaOH}$
- c. $\text{HC}_2\text{H}_3\text{O}_2 / \text{NH}_3$
- d. $\text{HNO}_2 / \text{Mg}(\text{OH})_2$
- e. $\text{HC}_2\text{H}_3\text{O}_2 / \text{KOH}$



AL CHEMISTRY II / جميع الشعب

How many grams Na (23.0 g/ mol) are present in 5.00 mL of 0.10 M solution Na_2SO_4 (142.1 g/ mol) ?

Select one:

- a. 0.041
- b. 0.023
- c. 0.051
- d. 0.060
- e. 0.032



GENERAL CHEMISTRY II / جميع الشعب

How many hydrogen atoms are there in 4.05 g of water?

Molar mass: $\text{H}_2\text{O} = 18.02 \text{ g/mol}$; Avogadro No. = 6.022×10^{23}

Select one:

- a.
 1.20×10^{22}
- b.
 2.69×10^{22}
- c.
 4.02×10^{22}
- d.
 2.71×10^{23}
- e.
 1.34×10^{23}



GENERAL CHEMISTRY II / جميع الشعب

30.0 mL of AgNO_3 solution ($\text{AgNO}_3 = 169.9 \text{ g/mol}$) of unknown concentration was found to give 1.50 grams of AgCl (143.4 g/mol) upon addition of excess sodium chloride. What is the molarity of the original AgNO_3 solution?

Select one:

- a. 0.262
- b. 0.174
- c. 0.349
- d. 0.523
- e. 0.209

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RAL CHEMISTRY II / جميع الشعب

Upon burning, a hydrocarbon composed of C and H gave 6.80 g of CO_2 and 1.39 g of H_2O .

What is the empirical formula of the hydrocarbon?

Molar mass: C = 12.0; H = 1.01; O = 16.0; CO_2 = 44.0; H_2O = 18.0

Select one:

- a. CH_2
- b. C_2H
- c. CH_4
- d. CH
- e. CH_3

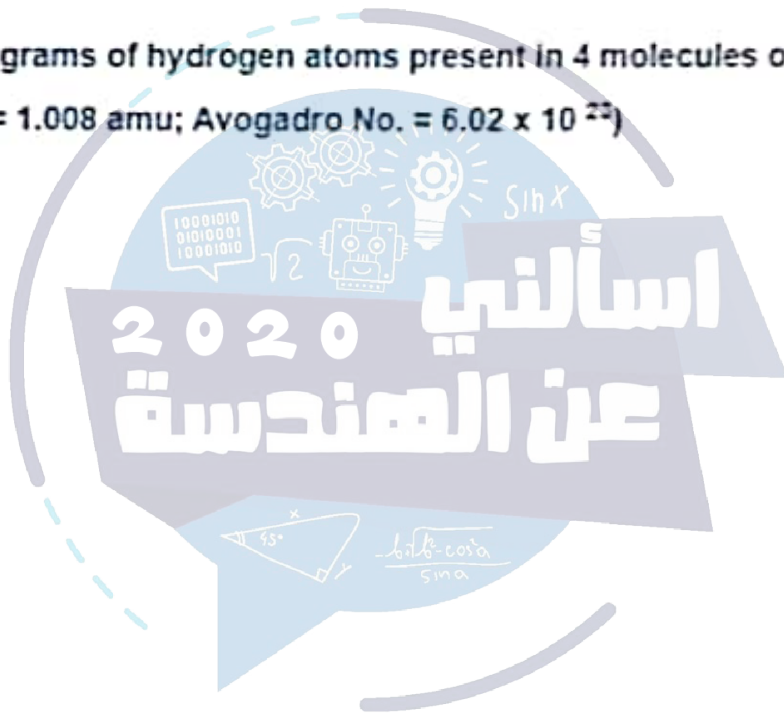
GENERAL CHEMISTRY II / جميع الشعب

What is the mass in grams of hydrogen atoms present in 4 molecules of water?

(Atomic mass of H = 1.008 amu; Avogadro No. = 6.02×10^{23})

Select one:

- a. 1.34×10^{-23}
- b. 2.01×10^{-23}
- c. 6.70×10^{-24}
- d. 1.00×10^{-23}
- e. 1.67×10^{-23}

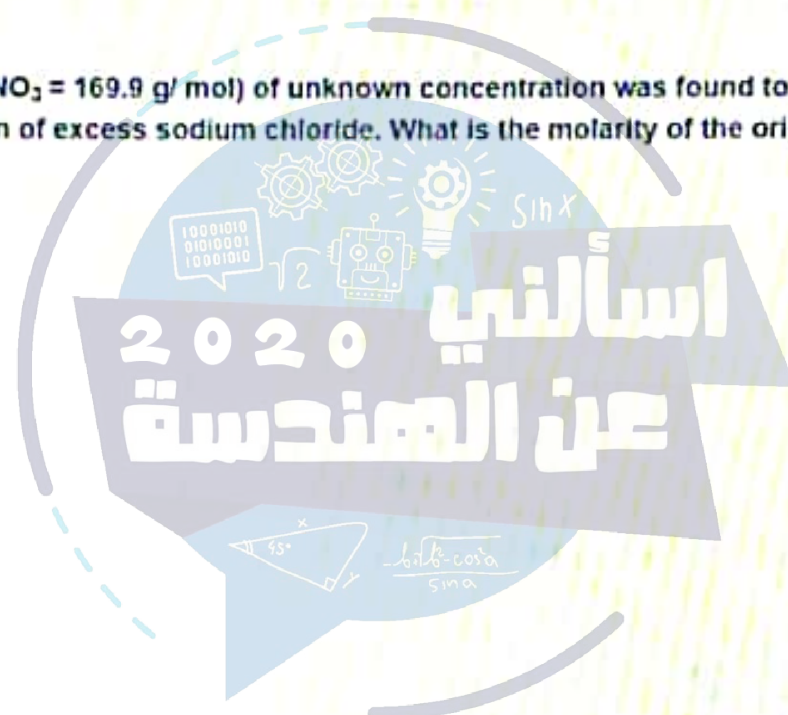


AL CHEMISTRY II / جميع الشعب

30.0 mL of AgNO_3 solution ($\text{AgNO}_3 = 169.9 \text{ g/mol}$) of unknown concentration was found to give 1.50 grams of AgCl (143.4 g/mol) upon addition of excess sodium chloride. What is the molarity of the original AgNO_3 solution?

Select one:

- a. 0.262
- b. 0.174
- c. 0.349
- d. 0.523
- e. 0.209



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GENERAL CHEMISTRY II / جميع الشعب

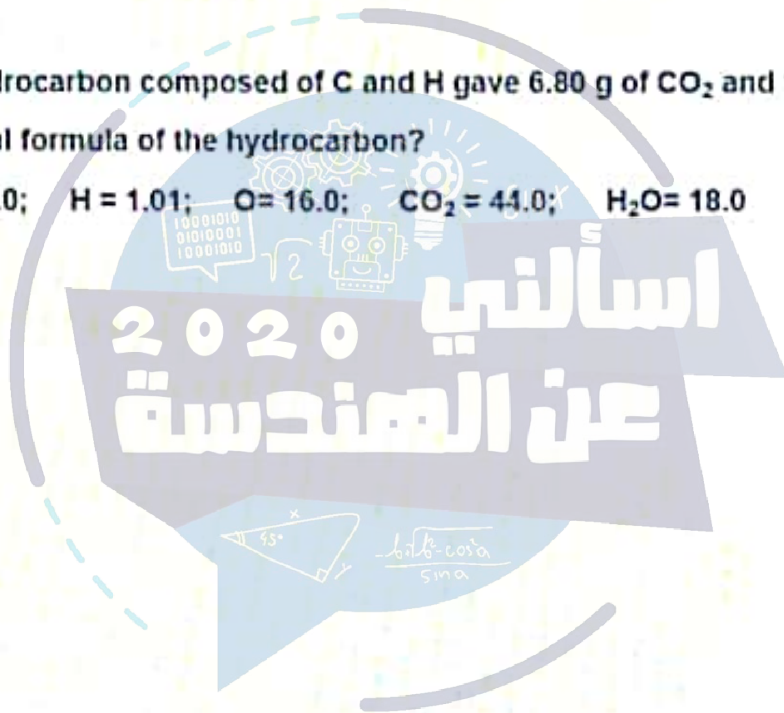
Upon burning, a hydrocarbon composed of C and H gave 6.80 g of CO_2 and 1.39 g of H_2O .

What is the empirical formula of the hydrocarbon?

Molar mass: C = 12.0; H = 1.01; O = 16.0; $\text{CO}_2 = 44.0$; $\text{H}_2\text{O} = 18.0$

Select one:

- a. CH_2
- b. C_2H
- c. CH_4
- d. CH
- e. CH_3



GENERAL CHEMISTRY II / جميع الشعب

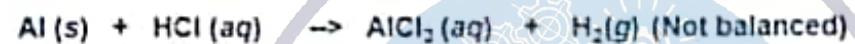
Question 6

What is the total number of electrons involved in the following oxidation-reduction reaction?

Marked out of 10

Flag question

What is the total number of electrons involved in the following oxidation-reduction reaction?



Select one:

- a. 6
- b. 2
- c. 8
- d. 10
- e. 4

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