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Chem. 101 Exams

Chem. 101 Final Exam

Question 3
Not answered
0 out of 10
0.00 %
0.00 %

The concentration of calcium in seawater sample was found to be 19000 mg/L. What is the percent concentration by mass of calcium in the sample assuming a density for seawater of 1.06 g/mL?

Select one:

- 0.230 %
- 0.905 %
- 0.0328 %
- 1.64 %
- 1.79 %

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Question 2

Not yet
answered

Marked out of
2.0

Flag
question

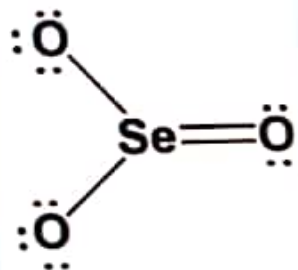
How can Te acquire a noble gas electron configuration?

Select one:

- By losing one electron
- By losing two electrons
- By gaining one electron
- By gaining two electrons
- By losing three electrons

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What formal charges are there on the following Lewis structure of SeO_3 ?



Select one:

- Se has +1, each O atom has 0
- Se has +1, two O atoms has -1 and one O atom has 0
- Se has +2, two O atoms has -1 and one O atom has 0
- Se has +2, one O atom has -2
- Se has 0, one O atom has +1

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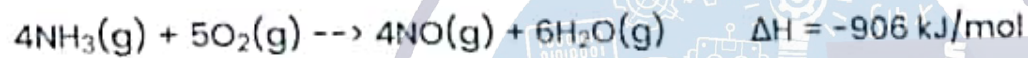
Question 4

Not yet answered

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Flag question

Calculate the change in internal energy (kJ/mol) when 4 moles of $\text{NH}_3(\text{g})$ are converted to 4 moles of $\text{NO}(\text{g})$ at 5 atm. and 425°C .



$R = 0.0821 \text{ Latm/mol.K}$ or 8.314 J/K.mol

Select one:

- 909.31
- 910.14
- 910.97
- 911.80
- 908.48

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Question 5

Not yet answered

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Flag question

How many electrons are there in the valence shell of Be in BeCl_2 ?

Select one:

- 6
- 10
- 4
- 12
- 14

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Not yet answered

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The average speed of nitrogen gas (N_2 , 28 g/mol) that effuses at $30.0\text{ }^\circ\text{C}$ is 480 m/s. The average speed at which butene gas (C_4H_8 , 56 g/mol) effuses at the same temperature is:

Select one:

- 481 m/s
- 339 m/s
- 354 m/s
- 495 m/s
- 396 m/s

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Question 7

Not yet answered

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Which one of the following solids have the largest melting point ?

Select one:

- RbI
- NaBr
- SrO
- CsCl
- KCl

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Question 8

Not yet answered

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Flag question

Which of the following atoms or ions are isoelectronic ?

Select one:

- K^+ and Cl
- Ca^{2+} and Mg^{2+}
- N^{3-} and F
- Li^+ and Be^{2+}
- Be^{2+} and B

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Question 9

Not yet answered

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Which of the following pair of elements would most likely form an ionic compound?

Select one:

- Ga and Mg
- Br and O
- Cl and I
- C and S
- Br and Rb

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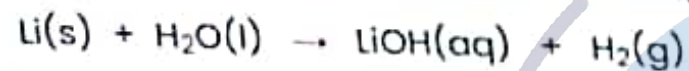
Question 10

Not yet answered

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Flag question

Lithium (Li) reacts with water according to the following unbalanced equation:



How many grams of Li are needed to produce 59.8g of H_2 ? (molar mass: H = 1.01, Li = 6.94 g/mol)

Select one:

- 411
- 273
- 136
- 342
- 205

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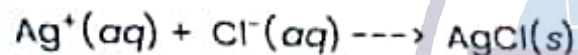
Question 11

Not yet answered

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How many milliliters of 0.165 M aluminum chloride (AlCl_3) are required to react completely with 35.0 mL of 0.210 M silver nitrate (AgNO_3)? The net ionic equation is:



Select one:

- 31.8 mL
- 19.1 mL
- 23.3 mL
- 14.8 mL
- 27.6 mL

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Question 14

Not yet answered

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In which of the following bonds would Br have a partial positive charge?

Select one:

- Br-Se
- Br-Ge
- Br-As
- Br-Cl
- Br-I

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Which one of the following electron configurations is considered a pseudo-Nobel gas configuration ?

Select one:

- [Kr]5s²4d⁷
- [Rn]5f³7s²6d¹
- [Xe]4f¹⁴6s²5d⁵
- [Rn]5f¹⁴
- [Ar]3d¹⁰4s²4p²

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Question 16

Not yet
answeredMarked out of
2.0Flag
question

What is the potential energy value (in kJ/mol) obtained via combining Cs^+ ions and F^- ions to form ionic bonds?

$$k = 8.99 \times 10^9 \text{ J.m/C}^2 \quad e = 1.6 \times 10^{-19} \text{ C}$$

$$\text{Avogadro No.} = 6.022 \times 10^{23}$$

The distance between ions = 0.350 nm

Select one:

- 462
- 554
- 513
- 396
- 693

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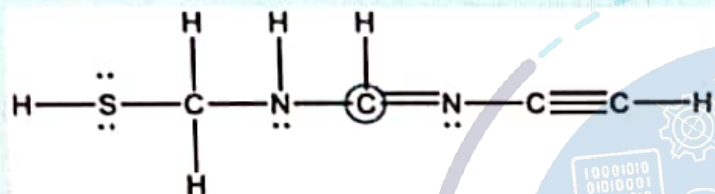
Question 15

Not yet answered

Marked out of 2.0

Flag question

Based on VSEPR, what is the expected geometry around the circled atom in the following compound?



Select one:

- Trigonal planar
- Bent
- Linear
- Tetrahedral
- Trigonal pyramidal

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Question 13

Not yet answered

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Flag question

The bond angles in SbI_5 are expected to be:

Select one:

- 180 °
- 90 °, 120 °, and 180 °
- 90 ° and 109.5 °
- 90 °
- 120 °

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Finish attempt...

Time left 0:57:06

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Question 12

Not yet answered

Marked out of 2.0

Flag question

What is the maximum number of electrons that can be placed in a shell with $n = 2$?

Select one:

- 10
- 2
- 6
- 8
- 14

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•

Which one of the following elements is expected to have the highest electronegativity?

Select one:

- Element (q): I.E = 4.0×10^3 kJ/mol and E.A = 2.0×10^3 kJ/mol
- Element (t) : I.E = 4.0×10^3 kJ/mol and E.A = 2.5×10^3 kJ/mol
- Element (z): I.E = 1.0×10^3 kJ/mol and E.A = 3.0×10^3 kJ/mol
- Element (y): I.E = 5.0×10^3 kJ/mol and E.A = 1.0×10^3 kJ/mol
- Element (x): I.E = 2.5×10^3 kJ/mol and E.A = 2.5×10^3 kJ/mol

Question 19

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Which of the following atoms is the most difficult to ionize?

Select one:

- Ne
- He
- F
- C
- O



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Ne

Question 20

Not yet answered

Marked out of 2.0

Flag question

Which one of the following statements is incorrect regarding electron affinity ?

Select one:

- In a given period, the electron affinity rises from the group 1A element to the group 7A element but with sharp drops in the group 2A and group 5A elements.
- Electron affinity is defined as the energy required to remove an electron from the atom's negative ion (in its ground state)
- Nobel gases have zero or small negative electron affinity values.
- The element with highest electron affinity in the periodic table is chlorine.
- Group 6A and group 7A elements have the smallest electron affinities of any of the other main-group elements

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Question 21

Not yet answered

Marked out of 2.0

Flag question

Which one of the following sets of quantum numbers is not possible?

Select one:

$n = 2, l = 1, m_l = 1$

$n = 1, l = 1, m_l = -1$

$n = 3, l = 2, m_l = -2$

$n = 2, l = 1, m_l = -1$

$n = 1, l = 0, m_l = 0$

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Question 22

Not yet answered

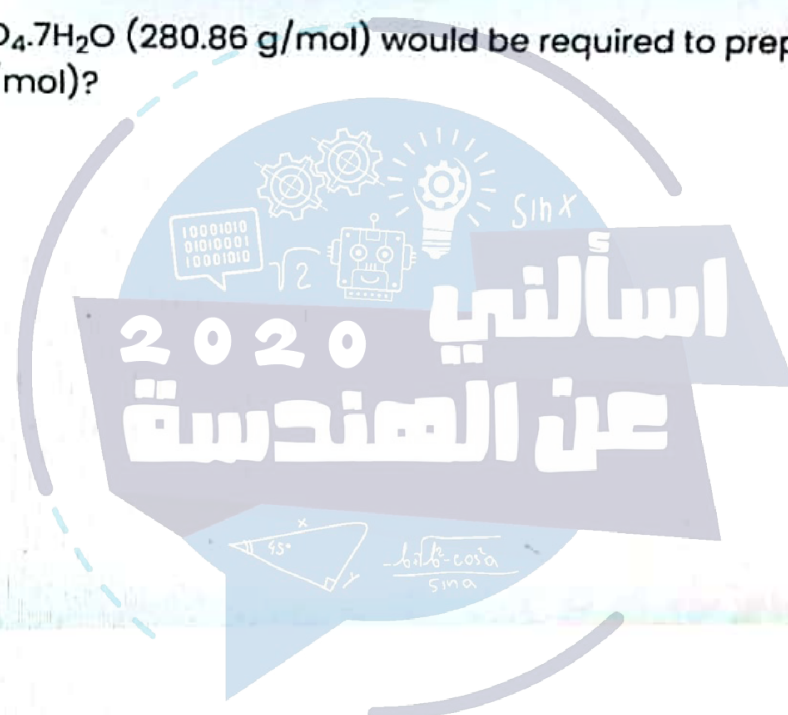
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Flag question

How many grams of $\text{NiSO}_4 \cdot 7\text{H}_2\text{O}$ (280.86 g/mol) would be required to prepare 1000 mL of a solution that is 0.300 M in $\text{NiSO}_4(aq)$ (154.76 g/mol)?

Select one:

- 42.1 g
- 50.6 g
- 59.0 g
- 67.4 g
- 84.3 g



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Question 23

Not yet
answered

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2.0

Flag
question

Which of the following ionic compounds does not exist ?

Select one:

- SrO
- BaI₂
- RbBr
- CsI
- CsO



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Question **24**

Not yet
answered

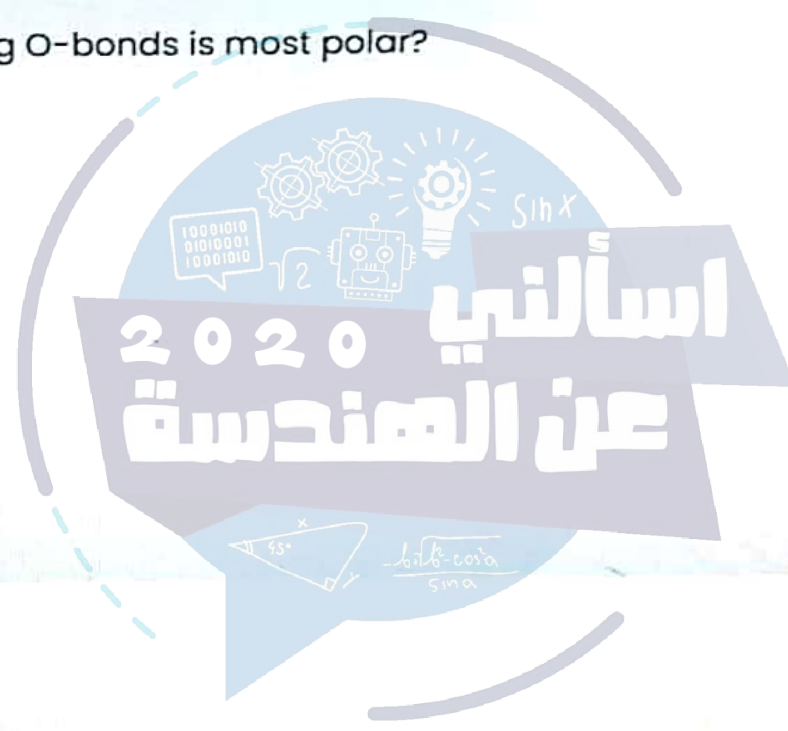
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2.0

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question

Which of the following O-bonds is most polar?

Select one:

- O-B
- O-C
- O-O
- O-F
- O-N



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Question 25

Not yet answered

Marked out of 2.0

Flag question

Which of the following atoms is smallest in size?

Select one:

- F
- B
- C
- Be
- O



Question 19

Not answered
Marked out of

Flag question

Which of the following atoms is the most difficult to ionize?

Select one:

- Ne
- He
- F
- C
- O



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Question 8

Not yet
answered

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2.0

Flag
question

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- Li^+ and Be^{2+}
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Question 16

Not yet answered

Marked out of 2.0

Flag question

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$$\text{Avogadro No.} = 6.022 \times 10^{23}$$

The distance between ions = 0.350 nm

Select one:

- 462
- 554
- 513
- 396
- 693



Question 18

Not yet answered

Marked out of 2.0

Flag question

Which one of the following elements is expected to have the highest electronegativity ?

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Question 20

Not yet answered

Marked out of 2.0

Flag question

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Question 13

Not yet answered

Marked out of 2.0

Flag question

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Select one:

- 180 °
- 90 °, 120 °, and 180 °
- 90 ° and 109.5 °
- 90 °
- 120 °

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Question **24**

Answer saved

Marked out of
2.0

🚩 Flag
question

Which of the following O-bonds is most polar?

Select one:

- O-B
- O-C
- O-O
- O-F
- O-N



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Question 16

Not yet answered

Marked out of 2.0

Flag question

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