Question 3

Not yet answered

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Flag question

The Food and Drug Administration (FDA) mandates that all vinegar sold must have concentration of 4% acetic acid by mass. Calculate the molarity of acetic acid in vinegar. Assume the density of vinegar is 1.00 g/ml. (Molar mass acetic acid= 60 g/mol)

- 0.24 M
- 2.40 M
- 0.66 M
- 0.06 M

[1] $4\% = \frac{M * 60}{1 * 10}\ \%$

$M = 0.67$

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Which one of the following statements is incorrect concerning acetic acid in vinegar experiment?

- Sodium hydroxide pellets can be handled by bare hands, but its dust is very irritating to the respiratory system
- It's not allowed to pipet solutions in the lab by mouth.
- KHP must be stored in the oven to keep moisture from adhering to the crystals
- When using 1.0 M NaOH solution to titrate an acid, the buret should be rinsed using 1.0 M NaOH solution just before the titration begins.
- When you run a titration using a buret, you can't start the titration if there are bubbles in it, even if you record the initial and final volumes.

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Which one of the following statements is correct concerning the acetic acid in vinegar and neutralizing capacity of antacid experiments?

- An antacid tablet is dissolved in H_2O first then titrated with standardized NaOH solution.
- One of the active ingredients in most commercial antacid tablets is $Ca(OH)_2$.
- The antacid tablet should be dissolved in an excess amount of HCl solution.
- The titration flask should be rinsed with vinegar solution before titration with NaOH solution.
- Starch can be used as an indicator in antacid experiment.

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